

# Ultrasonically Assisted Regioselective Nitration of Aromatic Compounds in Presence of Certain Group V and VI Metal Salts

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## ABSTRACT

Ultrasonically assisted nitration reactions (USANR) with anilides, moderately activated and non-activated aromatic compounds underwent smoothly and afforded good yields of products with high regio selectivity. Observed longer reaction times (6 - 8 hrs.) in metal catalyzed reactions reduced to (1 - 2 hrs.) under sonication. When ortho position is blocked para derivatives are obtained, and ortho nitro products are obtained when para position is blocked. In case of USANR of aromatic carbonyl and related compounds the effect of sonication is much more effective. The reactions could be completed only in few minutes.

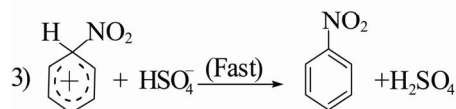
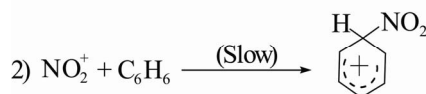
**Keywords:** Ultrasonically Assisted; Regioselective Nitration; Group V and VI Metal Salts; Reduction in Reaction Times

## 1. Introduction

Aromatic compounds such as benzene afford a variety of derivatives when they are allowed to react with other compounds under different conditions. The  $\pi$ -Electron cloud, located above and below the plane of benzene ring is more involved in holding carbon nuclei together through resonance than the  $\pi$ -electrons of a carbon-carbon double bond in an aliphatic unsaturated hydrocarbon. Yet, these  $\pi$ -electrons are much more loosely held in comparison with  $\pi$ -electrons and are easily available to a reagent that is seeking electrons. In its typical reactions benzene ring serves as a source of electrons. Benzene reacts with electron deficient molecules or species (electrophiles) in substitution reactions to afford benzene derivatives. A reaction is referred to as SE when a substitution reaction involves the attack by electrophiles [1-3].

Electrophilic nitration reactions are carried out by acid catalyzed reactions of  $\text{HNO}_3$  and its derivatives. The nitronium ion ( $\text{NO}_2^+$ ) is established as the reactive nitrating species by Ingold in 1940's following an early suggestion by Euler. Nitrating agents possess general formula  $\text{NO}_2\text{-X}$  which serve as sources of the  $\text{NO}_2^+$ , nitronium ion. The ease of X-elimination from  $\text{NO}_2\text{-X}$ , Euler derived a relative sequence of nitrating activity of certain nitrating agents: nitroniumion ( $\text{NO}_2^+$ ) > nitracidiumion

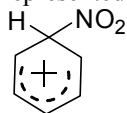
( $\text{NO}_2^+\text{-OH}_2$ ) > nitryl chloride ( $\text{NO}_2\text{-Cl}$ ) > nitrogen pentoxide ( $\text{N}_2\text{O}_5$ ) > methyl nitrate ( $\text{NO}_2\text{-O-CH}_3$ ). However, now the range of nitrating agents is much wider. The reactivity of the nitrating agent is also affected by the solvent used. A decreasing activity is observed in the series of solvents of increasing nucleophilicity: Sulfuric acid > nitric acid > nitro methane ( $\text{NO}_2\text{CH}_3$ ) > acetic acid > 1,4-dioxane > water. Nitration reaction proceeds through the formation of  $\text{NO}_2^+$  electrophile in the first step, followed by the formation of nitro derivative of carbocation in the second step. The carbocation thus formed supplements a proton to bisulphite ion to give  $\text{H}_2\text{SO}_4$  by affording nitro benzene.



This reaction is simply an acid-base equilibrium in which sulfuric acid serves as the acid and the much weaker nitric acid serves as a base. The very strong acid, sulfuric acid, causes nitric acid to ionize in the sense

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HO... NO<sub>2</sub><sup>+</sup>, rather than in the usual way of H<sup>+</sup>... -ONO<sub>2</sub>. The nitronium ion is well known, existing in salts such as nitronium perchlorate [NO<sub>2</sub><sup>+</sup> ClO<sub>4</sub><sup>-</sup>], and nitronium fluoroborate [NO<sub>2</sub><sup>+</sup> BF<sub>4</sub><sup>-</sup>]. Needing electrons, the nitronium ion finds them particularly available in the π-electron cloud of the Benzene ring and so in step [3-8], attaches itself to one of the carbon atoms by a covalent bond. This forms the carbocation; it's a resonance hybrid of the three structures, represented as:



Apart from therapeutic and diagnostic applications, *i.e.*, medical ultra sonography and teeth cleaning; ultrasound also has many applications in chemical and biochemical sciences. In the recent past the study of ultrasonically assisted reactions has been the subject of interest to synthetic and physical organic chemists probably due to a dramatic reduction in the reaction times followed by astonishingly high increase in the yields of reaction products [9,10]. The first example in which ultrasound induced a divergent pathway relative to thermal conditions was reported by Ando and co-workers more than 30 years ago [9-13]. Ever since the concept of chemical ultrasonics is introduced in 1927, a number of chemical reactions have been observed to occur under sonicated conditions [11-14]. One of the most fascinating areas of organic sonochemistry is sonochemical switching. In some cases, application of ultrasound may completely change the distribution of products or even cause the formation of different substances. Sonochemistry shares some aims with green chemistry, as it also uses smaller quantities of non-hazardous chemicals and solvents, reduces energy consumption, and increases product selectivity. The use of ultrasound in chemical reactions provides specific activation based on a physical phenomenon, known as acoustic cavitation. Recently, ultrasound has been utilized to accelerate a wide number of synthetically useful organic reactions [15-23]. In a recent review [20], David Bremner excellently elaborated some of the more modern uses of ultrasound in organic synthesis in different sections such as oxidations, carbon-nitrogen bond formation, carbon-carbon bond formation, cycloadditions, bond cleavage reactions, biological systems and sonochemical switching.

Since more than one decade our group is also actively working on exploiting the use of a variety of eco-friendly materials such as metal ions and surfactants as catalysts and non-conventional energy sources (such as microwave and ultra sound) to assist various organic transformations [24,25]. These reactions generally did not proceed at room temperature and is too sluggish even at elevated temperatures. However, under sonicated condi-

tions these reactions exhibited astonishingly high yields of the products and dramatic rate enhancements. The reaction times were reduced from several hours (in thermal reactions) to few minutes in ultrasonic irradiated reactions (the reaction flask was clamped in an ultrasonic bath of 2 MHz frequency at 40°C for the purpose of irradiations). Encouraged by these results we embarked on the use ultrasonics as a non-conventional energy source to assist nitration reactions of the present work.

## 2. Experimental Details

### 2.1. General Procedure

To a solution of 4-Hydroxy benzaldehyde 1b (0.099 ml, 1 mmol) in dichloromethane (DCM) or dichloro ethane (DCE) ammonium molybdate (1.235 gms, 1 mmol), and 70% HNO<sub>3</sub> (0.063 ml, 1 mmol) were added and the contents refluxed for 1hr. Progress of the reaction was monitored by TLC. After completion of the reaction, the reaction mixture was filtered and washed with water. The organic layer was separated and dried over Na<sub>2</sub>SO<sub>4</sub> and evaporated under vacuum. The crude product was purified by column chromatography using ethyl acetate-hexane (3:7) as eluent to get pure product 4-Hydroxy-3-nitro benzaldehyde as Yellow powder (mp 138°C - 140°C) in 85% yield. The nitrated products were characterized by comparing their spectral (IR, <sup>1</sup>H-NMR, mass, TLC) and physical data with authentic samples [26-31]. Certain representative spectra are given as supplementary data.

### 2.2. Spectroscopic Data

**4-Nitrophenol:** Yellow powder; mp 112°C - 114°C (lit. [26] 114°C); IR (KBr) ( $\nu_{\max}/\text{cm}^{-1}$ ): 3331, 1614, 1592, 1500, 1346; <sup>1</sup>H NMR (250 MHz, CDCl<sub>3</sub>):  $\delta$  8.18 (d, 2H, *J* = 5.50 Hz, Ar-*H*), 7.00 (d, 2H, *J* = 5.49 Hz, Ar-*H*), 6.35 (s, 1H, -OH); GC-MS/EI: *m/z* (%) = 139 (M<sup>+</sup>, 14), 109 (50), 81 (33), 65 (84), 53 (31), 46 (22), 39 (100), 30 (77), 28 (41).

**2-Nitrophenol:** Yellow powder; mp 42°C - 44°C (lit. [26] 45°C); IR (KBr) ( $\nu_{\max}/\text{cm}^{-1}$ ): 3241, 1623, 1606, 1593, 1538, 1479; <sup>1</sup>H NMR (250 MHz, CDCl<sub>3</sub>):  $\delta$  10.5 (s, 1H, -OH), 8.10 (d, 1H, *J* = 8.50 Hz, Ar-*H*), 7.59 (m, 1H, Ar-*H*), 7.16 (d, 1H, *J* = 8.4 Hz, Ar-*H*), 7.00 (m, 1H, Ar-*H*); GC-MS/ EI: *m/z* (%) = 139 (M<sup>+</sup>, 40), 109 (35), 81 (55), 65 (60), 53 (38), 39 (100), 30 (66).

**4-Chloro-2-nitrophenol:** Yellow powder; mp 84°C - 86°C (lit [27] 85°C - 89°C); IR (KBr) ( $\nu_{\max}/\text{cm}^{-1}$ ): 3490, 1631, 1531, 1469, 1384, 1346, 1238; <sup>1</sup>H NMR (250 MHz, DMSO): 11.33 (brs, 1H, -OH), 7.90 (s, 1H, Ar-*H*), 7.54 (d, 1H, *J* = 8.9 Hz), 7.14 (d, 1H, *J* = 8.9 Hz, Ar-*H*); GC-MS/ EI: *m/z* (%) = 173 (M<sup>+</sup>, 16), 156 (7), 143 (26), 128 (10), 115 (41), 99 (39), 73 (29), 63 (100), 51 (20), 37 (13), 30 (26).

**4-Chloro-3-nitrobenzaldehyde:** Yellow powder; mp

140°C - 147°C (lit. [31] 148°C - 150°C); IR (KBr) ( $\nu_{\max}/\text{cm}^{-1}$ ): 3101, 2873, 2715, 2596, 1701, 1593, 1431, 1330; <sup>1</sup>H NMR (250 MHz, DMSO):  $\delta$  10.1 (s, 1H, -CHO), 8.22 (s, 1H, Ar-H), 7.90 (d, 1H,  $J = 3.1$  Hz, Ar-H), 7.55 (d, 1H,  $J = 3.1$  Hz, Ar-H); GC-MS/ EI:  $m/z$  (%) = 185 (M<sup>+</sup>, 9), 167 (10), 149 (22), 129 (15), 97 (19), 73 (45), 57 (100).

**2,4-Dinitrophenol:** Yellow powder; mp 107°C - 110°C (lit. [31] 108°C - 112°C); IR (KBr) ( $\nu_{\max}/\text{cm}^{-1}$ ): 3587, 3529, 1608, 1569, 1531, 1523, 1438, 1357, 1280; <sup>1</sup>H NMR (250 MHz, DMSO):  $\delta$  11.02 (brs, 1H, -OH), 8.65 (s, 1H, Ar-H), 8.26 (d, 1H,  $J = 9.3$  Hz), 7.22 (d, 1H,  $J = 9.2$  Hz, Ar-H); GC-MS/ EI:  $m/z$  (%) = 184 (M<sup>+</sup>, 4), 167 (10), 129 (10), 111 (12), 85 (22), 69 (100).

**4-Hydroxy-3-nitrobenzaldehyde:** Yellow powder; mp 138°C - 140°C (lit [31] 140°C - 141°C); IR (KBr) ( $\nu_{\max}/\text{cm}^{-1}$ ): 3236, 2816, 1693, 1620, 1562, 1423, 1334, 1269, 1126; <sup>1</sup>H NMR (250 MHz, CDCl<sub>3</sub>):  $\delta$  11.01 (brs, 1H, -OH), 9.94 (s, 1H, -CHO), 8.63 (s, 1H, Ar-H), 8.14 (d, 1H,  $J = 3.5$  Hz), 7.33 (d, 1H,  $J = 3.5$  Hz, Ar-H); GC-MS/ EI:  $m/z$  (%) = 167 (M<sup>+</sup>, 10), 149 (25), 129 (11), 111 (11), 85 (23), 69 (100).

**2-Hydroxy-5-nitrobenzaldehyde:** Yellow powder; mp 127°C - 129°C (lit. [27] 129°C - 130°C); IR (KBr) ( $\nu_{\max}/\text{cm}^{-1}$ ): 3460, 3074, 2935, 1666, 1620, 1477, 1334, 1323; <sup>1</sup>H NMR (250 MHz, DMSO):  $\delta$  11.5 (brs, 1H, -OH), 9.91 (s, 1H, -CHO), 8.49 (s, 1H, Ar-H), 8.32 (d, 1H,  $J = 9.1$  Hz, Ar-H), 7.05 (d, 1H,  $J = 9.1$  Hz, Ar-H); GC-MS/EI:  $m/z$  (%) = 167 (M<sup>+</sup>, 11), 149 (13), 129 (7), 111 (12), 85 (24), 69 (100).

**4-Nitroaniline:** Yellow powder; mp (146°C - 148°C (lit. [27] 146°C - 149°C); IR (KBr) ( $\nu_{\max}/\text{cm}^{-1}$ ): 3475, 3363, 1601, 1530, 1348; <sup>1</sup>H NMR (250 MHz, CDCl<sub>3</sub>):  $\delta$  7.94 (d, 2H,  $J = 5.50$  Hz, Ar-H), 6.76 (d, 2H,  $J = 5.49$  Hz, Ar-H), 4.25 (brs, 2H, -NH<sub>2</sub>); GC-MS/ EI:  $m/z$  (%) = 138 (M<sup>+</sup>, 45).

**3-Hydroxy-4-nitro Acetophenone:**  $\delta$  10.68 (s 1H, OH) 2.38 (s 3H, COMe) 7.62 (d, 1H,  $J = 8.5$ Hz), 8.10 (m 1H,  $J = 8.5$  Hz,  $J = 3.5$ Hz) 7.52 (d 1H,  $J = 3.5$ Hz);  $m/z = 181$ .

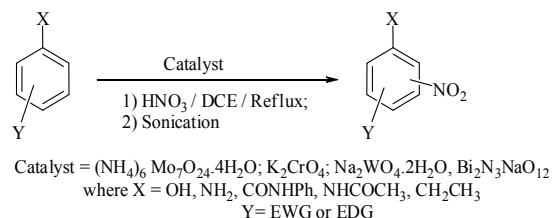
**4-Amino-2-nitro Phenol:**  $\delta$  4.45 (s 2H, NH<sub>2</sub>) 9.92 (s 1H, OH) 6.76 (d, 1H,  $J = 8$ Hz) 6.82 (d 1H,  $J = 8$ Hz), 7.24 (s 1H);  $m/z = 154$ .

### 3. Results & Discussion

#### 3.1. Nitration Reactions under Conventional (Thermal) Conditions

Nitration of non-active and moderately active aromatic compounds were catalyzed by Bismuth nitrate (BN), sodium Bismuthate (SB) and Group VI B metal salts such as potassium chromate (PCR), ammonium molybdate (AMB) and sodium tungstate (STG). To substantiate the generality, the reaction is conducted with an array of non-activated, moderately activated and activated organic compounds. General reaction conditions are shown in **Scheme 1**.

The data presented in **Tables 1-5** show that the reactions with aromatic compounds such as phenols (entries 1 to 4) anilides (entries 7 to 14) underwent smoothly with

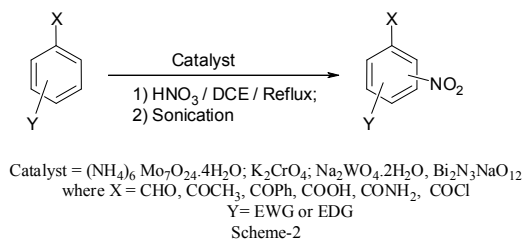


**Scheme 1. Nitration of aromatic compounds.**

moderate to long reaction times (6 to 8 hours) and afforded good yields of products with good regioselectivity. It is of interest to note that para substituted compounds (entries 2 to 4, 8, 9, 11 to 14 in **Tables 1-5**) underwent nitration only at ortho position and afforded excellent yield of products, while ortho substituted compound (entry 7) gave very good yield of para derivative. Under similar conditions aniline, chloro benzene, toluene and ethyl benzene (entries 5, 19, 20, 21) have undergone smooth nitration with high regioselectivity.

Even though, benzene failed to undergo nitration; it is also of interest to note that aromatic hydrocarbons such as anthracene, naphthalene,  $\alpha$ -naphthol,  $\beta$ -naphthol, 2-amino anthracene and 1-naphthalde-hyde underwent smooth nitration with quantitative yields (80% to 90%). When the reaction protocol is extended to activated-aromatic compounds such as carbonyl compounds, interestingly the reaction times are drastically reduced from several hours to about one hour. Related data are given in **Tables 6-10**.

Nitration of aromatic carbonyl and related compounds underwent rapidly affording high yields of the corresponding mono nitro derivatives with high regioselectivity. The reactions were clean, no attack being observed on the alkyl portion of the ketones. The present nitration produces a remarkably high yield of para nitro compounds (entries 1a, 1c, 1g, 1h, 1j, 1k, 1l, 1m, 1o, 1p & 1q). To check the regioselectivity of the reaction, the reaction is carried out with different para substituted aromatic carbonyl and related compounds (entries 1b, 1d, 1e, 1f, 1i, 1r) which afforded only meta nitrated product in good yield. General reaction conditions are given in **Scheme 2**. In the absence of metal salts, the nitration did not proceed. The results of the present work are summarized in **Tables 6-10**.



**Scheme 2. Nitration of aromatic aldehydes.**

**Table 1. Ammonium molybdate as catalyst for regio selective nitration of anilides, non-activated and moderately activated organic compounds under mild acid conditions.**

S.N.	Substrate	Conventional			USANR		
		Time (hr)	Yield (%)		Time (min)	Yield (%)	
			Para	Ortho		Para	Ortho
1	Phenol	8	80	10	80	82	09
2	4-Chloro phenol	8	-	85	80	-	81
3	4- Nitro phenol	7	-	84	70	-	80
4	4-Amino phenol	7	-	80	70	-	79
5	Aniline	8	74	20	80	78	11
6	Acetanilide	8	90	-	80	92	-
7	2-Chloro acetanilide	6	100	-	60	92	-
8	4-Chloro acetanilide	6	-	99	60	-	92
9	4-Nitro acetanilide	6	-	97	60	-	98
10	3-Nitro acetanilide	6	65	29	60	68	25
11	4-Methyl acetanilide	6	-	92	60	-	92
12	4-Flouro acetanilide	6	-	96	60	-	96
13	4-Bromo acetanilide	6	-	94	60	-	95
14	4-Hydroxy acetanilide	6	-	87	60	-	85
15	Benzanilide	6	86	12	60	88	10
16	2-Chloro benzanilide	6	86	-	60	88	-
17	4-Chloro benzanilide	6	-	92	60	-	90
18	4-Nitro benzanilide	6	-	88	60	-	90
19	Chloro benzene	6	82	16	60	84	11
20	Toluene	6	74	20	60	76	12
21	Ethyl benzene	6	83	12	60	85	09

**Table 2. Ultrasonically assisted potassium chromate catalyzed regio selective nitration of anilides, non-activated and moderately activated organic compounds under mild acid conditions.**

S.N.	Substrate	Conventional			USANR		
		Time (hr)	Yield (%)		Time (min)	Yield (%)	
			Para	Ortho		Para	Ortho
1	Phenol	9	78	09	90	82	12
2	4-Chloro phenol	9	-	81	90	-	88
3	4- Nitro phenol	8	-	80	80	-	89
4	4-Amino phenol	8	-	79	80	-	86
5	Aniline	9	83	11	90	80	11
6	Acetanilide	9	86	-	90	90	-
7	2-Chloro acetanilide	7	90	-	70	99	-
8	4-Chloro acetanilide	6	-	90	60	-	98
9	4-Nitro acetanilide	7	-	88	70	-	97
10	3-Nitro acetanilide	7	65	25	70	66	25
11	4-Methyl acetanilide	7	-	86	70	-	94
12	4-Flouro acetanilide	7	-	89	70	-	97
13	4-Bromo acetanilide	7	-	85	70	-	96
14	4-Hydroxy acetanilide	7	-	81	70	-	89
15	Benzanilide	7	83	14	70	87	08
16	2-Chloro benzanilide	7	82	-	70	88	-
17	4-Chloro benzanilide	7	-	90	70	-	94
18	4-Nitro benzanilide	7	-	86	70	-	90
19	Chloro benzene	7	80	11	70	83	14
20	Toluene	7	70	16	70	76	18
21	Ethyl benzene	7	80	09	70	85	10

**Table 3. Ultrasonically assisted sodium tungstate catalyzed regio selective nitration of anilides, non-activated and moderately activated organic compounds under mild acid conditions.**

S.N.	Substrate	Conventional			USANR		
		Time (hr)	Yield (%)		Time (min)	Yield (%)	
			Para	Ortho		Para	Ortho
1	Phenol	7	82	12	70	84	12
2	4-Chloro phenol	7	-	88	70	-	89
3	4- Nitro phenol	6	-	89	60	-	90
4	4-Amino phenol	6	-	86	60	-	86
5	Aniline	7	80	11	70	82	11
6	Acetanilide	7	90	-	70	92	-
7	2-Chloro acetanilide	5	99	-	50	99	-
8	4-Chloro acetanilide	5	-	98	50	-	98
9	4-Nitro acetanilide	5	-	97	50	-	97
10	3-Nitro acetanilide	6	66	25	60	68	20
11	4-Methyl acetanilide	5	-	94	50	-	95
12	4-Flouro acetanilide	5	-	97	50	-	97
13	4-Bromo acetanilide	5	-	96	50	-	98
14	4-Hydroxy acetanilide	6	-	89	60	-	90
15	Benzanilide	5	87	08	50	88	08
16	2-Chloro benzanilide	6	88	-	60	90	-
17	4-Chloro benzanilide	5	-	94	50	-	95
18	4-Nitro benzanilide	6	-	90	60	-	91
19	Chloro benzene	6	83	14	60	84	12
20	Toluene	6	76	18	60	78	14
21	Ethyl benzene	5	85	10	50	86	10

**Table 4. Ultrasonically assisted bismuth nitrate catalyzed nitration of non-active and moderately active aromatic organic compounds under mild acid conditions.**

S.N.	Substrate	Conventional			USANR		
		Time (hr)	Yield (%)		Time (min)	Yield (%)	
			Para	Ortho		Para	Ortho
1	Phenol	9	78	09	90	82	12
2	4-Chloro phenol	9	-	81	90	-	88
3	4- Nitro phenol	8	-	80	80	-	89
4	4-Amino phenol	8	-	79	80	-	86
5	Aniline	9	83	11	90	85	11
6	Acetanilide	9	86	-	90	90	-
7	2-Chloro acetanilide	7	90	-	70	99	-
8	4-Chloro acetanilide	6	-	90	60	-	98
9	4-Nitro acetanilide	7	-	88	70	-	97
10	3-Nitro acetanilide	7	65	25	70	66	25
11	4-Methyl acetanilide	7	-	86	70	-	94
12	4-Flouro acetanilide	7	-	89	70	-	97
13	4-Bromo acetanilide	7	-	85	70	-	96
14	4-Hydroxy acetanilide	7	-	81	70	-	89
15	Benzanilide	7	83	14	70	87	08
16	2-Chloro benzanilide	7	82	-	70	88	-
17	4-Chloro benzanilide	7	-	90	70	-	94
18	4-Nitro benzanilide	7	-	86	70	-	90
19	Chloro benzene	7	80	11	70	83	14
20	Toluene	7	70	16	70	76	18
21	Ethyl benzene	7	80	09	70	85	10

**Table 5. Ultrasonically assisted sodium bismuthate catalyzed nitration of non-active and moderately active aromatic organic compounds under mild acid conditions.**

S.N.	Substrate	Conventional			USANR		
		Time (hr)	Yield (%)		Time (min)	Yield (%)	
			Para	Ortho		Para	Ortho
1	Phenol	7	82	12	70	85	12
2	4-Chloro phenol	7	-	88	70	-	88
3	4- Nitro phenol	6	-	89	60	-	89
4	4-Amino phenol	6	-	86	60	-	86
5	Aniline	7	80	11	70	81	11
6	Acetanilide	7	90	-	70	92	-
7	2-Chloro acetanilide	5	99	-	50	99	-
8	4-Chloro acetanilide	5	-	98	50	-	98
9	4-Nitro acetanilide	5	-	97	50	-	98
10	3-Nitro acetanilide	6	66	20	60	68	20
11	4-Methyl acetanilide	5	-	94	50	-	95
12	4-Flouro acetanilide	5	-	97	50	-	98
13	4-Bromo acetanilide	5	-	96	50	-	97
14	4-Hydroxy acetanilide	6	-	89	60	-	89
15	Benzanilide	5	87	08	50	89	08
16	2-Chloro benzanilide	6	88	-	60	89	-
17	4-Chloro benzanilide	5	-	94	50	-	95
18	4-Nitro benzanilide	6	-	90	60	-	91
19	Chloro benzene	6	83	14	60	85	12
20	Toluene	6	76	18	60	78	16
21	Ethyl benzene	5	85	10	50	86	10

**Table 6. Ultrasonically assisted potassium chromate catalyzed nitration of carbonyl and related compounds under mild acid conditions.**

Entry	Substrate	Product	Conventional	USANR
			R.T (100 min)	R.T (75 min)
			Yield (%)	Yield (%)
1a	Benzaldehyde	4-Nitro benzaldehyde	81	88
1b	4-Hydroxy benzaldehyde	4-Hydroxy-3-nitro benzaldehyde	80	86
1c	2,6-Dichloro benzaldehyde	2,6-Dichloro-4-nitro benzaldehyde	79	84
1d	4-Chloro benzaldehyde	4-Chloro-3-nitro benzaldehyde	80	89
1e	Salicylaldehyde	2-Hydroxy-5-nitro benzaldehyde	76	79
1f	3,4-Dimethoxy benzaldehyde	3,4-Dimethoxy-5-nitro-benzaldehyde	80	89
1g	Acetophenone	4-Nitro acetophenone	77	86
1h	Benzophenone	4-Nitro benzophenone	80	87
1i	4-Hydroxy acetophenone	4-Hydroxy-3-nitro acetophenone	79	88
1j	2,4-Dihydroxy acetophenone	5-Nitro-2,4-dihydroxy acetophenone	78	86
1k	2-Amino benzophenone	2-Amino-5-nitro benzophenone	83	89
1l	Benzoic acid	4-Nitro benzoic acid	81	89
1m	2-Chlorobenzoic acid	2-Chloro-4-nitro benzoic acid	82	91
1n	Salicylic acid	2-Hydroxy-5-nitro benzoic acid	80	90
1o	Benzoyl chloride	4-Nitrobenzoyl chloride	80	89
1p	Methylbenzoate	4-Nitromethyl benzoate	83	90
1q	Benzamide	4-Nitro benzamide	81	88
1r	p-Toluene sulphonic acid	3-Nitro-p-toluene sulphonic acid	80	87
1s	Nitrobenzene	1,3-Dinitro benzene	82	86

**Table 7. Ultrasonically assisted ammonium molybdate catalyzed nitration of carbonyl and related compounds under mild acid conditions.**

Entry	Substrate	Product	Conventional	USANR
			R.T (70 min)	R.T (40 min)
			Yield (%)	Yield (%)
1a	Benzaldehyde	4-Nitro benzaldehyde	85	88
1b	4-Hydroxy benzaldehyde	4-Hydroxy-3-nitro benzaldehyde	84	86
1c	2,6-Dichloro benzaldehyde	2,6-Dichloro-4-nitro benzaldehyde	82	84
1d	4-Chloro benzaldehyde	4-Chloro-3-nitro benzaldehyde	86	89
1e	Salicylaldehyde	2-Hydroxy-5-nitro benzaldehyde	78	80
1f	3,4-Dimethoxy benzaldehyde	3,4-Dimethoxy-5-nitro- benzaldehyde	84	89
1g	Acetophenone	4-Nitro acetophenone	82	86
1h	Benzophenone	4-Nitro benzophenone	82	87
1i	4-Hydroxy acetophenone	4-Hydroxy-3-nitro acetophenone	82	88
1j	2,4-Dihydroxy acetophenone	5-Nitro-2,4-dihydroxy acetophenone	80	86
1k	2-Amino benzophenone	2-Amino-5-nitro benzophenone	84	89
1l	Benzoic acid	4-Nitro benzoic acid	86	89
1m	2-Chlorobenzoic acid	2-Chloro-4-nitro benzoic acid	88	91
1n	Salicylic acid	2-Hydroxy-5-nitro benzoic acid	82	90
1o	Benzoyl chloride	4-Nitrobenzoyl chloride	84	89
1p	Methylbenzoate	4-Nitromethyl benzoate	86	90
1q	Benzamide	4-Nitro benzamide	84	88
1r	p-Toluene sulphonic acid	3-Nitro-p-toluene sulphonic acid	82	87
1s	Nitrobenzene	1,3-Dinitro benzene	84	89

**Table 8. Ultrasonically assisted sodium tungstate catalyzed nitration of carbonyl and related compounds under mild acid conditions.**

Entry	Substrate	Product	Conventional	USANR
			R.T (70 min)	R.T (40 min)
			Yield (%)	Yield (%)
1a	Benzaldehyde	4-Nitro benzaldehyde	88	89
1b	4-Hydroxy benzaldehyde	4-Hydroxy-3-nitro benzaldehyde	86	90
1c	2,6-Dichloro benzaldehyde	2,6-Dichloro-4-nitro benzaldehyde	84	85
1d	4-Chloro benzaldehyde	4-Chloro-3-nitro benzaldehyde	89	91
1e	Salicylaldehyde	2-Hydroxy-5-nitro benzaldehyde	79	80
1f	3,4-Dimethoxy benzaldehyde	3,4-Dimethoxy-5-nitro-benzaldehyde	89	91
1g	Acetophenone	4-Nitro acetophenone	86	88
1h	Benzophenone	4-Nitro benzophenone	87	89
1i	4-Hydroxy acetophenone	4-Hydroxy-3-nitro acetophenone	88	90
1j	2,4-Dihydroxy acetophenone	5-Nitro-2,4-dihydroxy acetophenone	86	89
1k	2-Amino benzophenone	2-Amino-5-nitro benzophenone	89	90
1l	Benzoic acid	4-Nitro benzoic acid	89	91
1m	2-Chlorobenzoic acid	2-Chloro-4-nitro benzoic acid	91	92
1n	Salicylic acid	2-Hydroxy-5-nitro benzoic acid	90	92
1o	Benzoyl chloride	4-Nitrobenzoyl chloride	89	91
1p	Methylbenzoate	4-Nitromethyl benzoate	90	91
1q	Benzamide	4-Nitro benzamide	88	90
1r	p-Toluene sulphonic acid	3-Nitro-p-toluene sulphonic acid	87	88
1s	Nitrobenzene	1,3-Dinitro benzene	86	89

**Table 9. Ultrasonically assisted sodium bismuthate catalyzed nitration of carbonyl and related compounds under mild acid conditions.**

Entry	Substrate	Product	Conventional	USANR
			R.T (75 min)	R.T (40 min)
			Yield (%)	Yield (%)
1a	Benzaldehyde	4-Nitro benzaldehyde	81	85
1b	4-Hydroxy benzaldehyde	4-Hydroxy-3-nitro benzaldehyde	80	82
1c	2,6-Dichloro benzaldehyde	2,6-Dichloro-4-nitro benzaldehyde	79	80
1d	4-Chloro benzaldehyde	4-Chloro-3-nitro benzaldehyde	80	89
1e	Salicylaldehyde	2-Hydroxy-5-nitro benzaldehyde	76	79
1f	3,4-Dimethoxy Benzaldehyde	3,4-Dimethoxy-5-nitro- benzaldehyde	80	85
1g	Acetophenone	4-Nitro acetophenone	77	86
1h	Benzophenone	4-Nitro benzophenone	80	87
1i	4-Hydroxy acetophenone	4-Hydroxy-3-nitro acetophenone	79	88
1j	2,4-Dihydroxy acetophenone	5-Nitro-2,4-dihydroxy acetophenone	78	86
1k	2-Amino benzophenone	2-Amino-5-nitro benzophenone	83	89
1l	Benzoic acid	4-Nitro benzoic acid	81	89
1m	2-Chlorobenzoic acid	2-Chloro-4-nitro benzoic acid	82	91
1n	Salicylic acid	2-Hydroxy-5-nitro benzoic acid	80	90
1o	Benzoyl chloride	4-Nitrobenzoyl chloride	80	89
1p	Methylbenzoate	4-Nitromethyl benzoate	83	90
1q	Benzamide	4-Nitro benzamide	81	88
1r	p-Toluene sulphonic acid	3-Nitro-p-toluene sulphonic acid	80	87
1s	Nitrobenzene	1,3-Dinitro benzene	82	86

**Table 10. Ultrasonically assisted bismuth nitrate catalyzed nitration of carbonyl and related compounds under mild acid conditions.**

Entry	Substrate	Product	Conventional	USANR
			R.T (90 min)	R.T (60 min)
			Yield (%)	Yield (%)
1a	Benzaldehyde	4-Nitro benzaldehyde	88	91
1b	4-Hydroxy benzaldehyde	4-Hydroxy-3-nitro benzaldehyde	86	87
1c	2,6-Dichloro benzaldehyde	2,6-Dichloro-4-nitro benzaldehyde	84	86
1d	4-Chloro benzaldehyde	4-Chloro-3-nitro benzaldehyde	89	92
1e	Salicylaldehyde	2-Hydroxy-5-nitro benzaldehyde	79	81
1f	3,4-Dimethoxy Benzaldehyde	3,4-Dimethoxy-5-nitro- benzaldehyde	89	90
1g	Acetophenone	4-Nitro acetophenone	86	88
1h	Benzophenone	4-Nitro benzophenone	87	89
1i	4-Hydroxy acetophenone	4-Hydroxy-3-nitro acetophenone	88	90
1j	2,4-Dihydroxy acetophenone	5-Nitro-2,4-dihydroxy acetophenone	86	89
1k	2-Amino benzophenone	2-Amino-5-nitro benzophenone	89	91
1l	Benzoic acid	4-Nitro benzoic acid	89	90
1m	2-Chlorobenzoic acid	2-Chloro-4-nitro benzoic acid	91	92
1n	Salicylic acid	2-Hydroxy-5-nitro benzoic acid	90	91
1o	Benzoyl chloride	4-Nitrobenzoyl chloride	89	91
1p	Methylbenzoate	4-Nitromethyl benzoate	90	92
1q	Benzamide	4-Nitro benzamide	88	89
1r	p-Toluene sulphonic acid	3-Nitro-p-toluene sulphonic acid	87	88
1s	Nitrobenzene	1,3-Dinitro benzene	86	89



The observed para selectivity could be explained due to the reversible nucleophilic hydration of carbonyl group by utilizing the water molecules present in the molybdenum species. Hydration of the carbonyl group makes it a non-electron withdrawing species, which in turn facilitates para nitration in the subsequent steps. Further, it is of great interest to note that benzoyl chloride (entry 10) also underwent para nitration without loss of chloride. It is also found that in no case under these conditions oxidation of aldehydic group to corresponding carboxylic group has been observed. The results clearly indicate that the direct nitration of aromatic carbonyl and related compounds can be achieved successfully by employing ammonium molybdate and nitric acid as nitrating agent with high para-regioselectivity.

### 3.2. Ultrasonically Assisted Nitration Reactions

Certain recent non-conventional activation techniques, such as ultrasonically assisted methods have become stimulus to physical and organic chemists because of their use to enhance and promote organic synthesis. Ultrasonically assisted organic reactions (USAR) afforded dramatic rate enhancements followed by achieving the selectivity, ease of experimental manipulation. The use of this non-traditional tool has been found to of great help in overcoming many of the difficulties associated with conventional reactions, and offered both processes related and environmental advantages.

In the present study we have conducted all the nitration reactions under sonication. All the reactions in the present study underwent with highly remarkable rate and yield enhancements followed by a reduction in the reaction times as could be seen from the data presented in **Tables 1-10**.

The spectrum of ultrasound spreads over a wide range of frequencies, (10 Hz - 10 MHz) as shown in **Figure 1**. A higher frequency range (1 MHz -10 MHz) is used for medical and therapeutic purposes, medium frequency range (20 kHz - 100 kHz) is used for sonochemistry and low frequency range is used for human handling (**Figure 1**).

Literature reports have demonstrated that ultrasonic irradiations not only accelerate chemical reactions, but also promote new systems and also reduce the number of

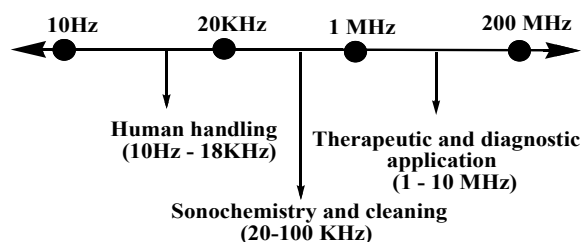


Figure 1. Ultrasound range diagram.

steps which are generally required for normal reactions. It is also well documented in literature that the ultrasonically assisted reactions occur through a transient state of high temperature and pressure and the chemical effects of ultrasound are due to the phenomenon of acoustic cavitation [12,13], as shown in **Figure 2**. When the reaction takes under ultrasonic conditions, sound waves pass through the solution containing reaction mixture and induce molecular motion through a series of compression and rarefaction cycles. At sufficiently high powers, the rarefaction cycle is able to pull molecules apart and create voids that are effectively bubbles containing small quantities of solvent vapour. At ultrasonic frequencies the succeeding compression wave does not fully collapse the bubbles and they grow in size through subsequent cycles by a process known as rectified diffusion. Eventually they become unstable and undergo violent collapse in a process known as acoustic cavitation, as shown in **Figure 2** [14,15]. Depending on the frequency, extreme temperatures and pressures can be generated within the bubble and at the point of bubble collapse. This phenomenon provides a source of energy which can be used to enhance the rate of a chemical process. Each cavitation bubble acts itself as a localized micro reactor which causes rate accelerations in the ultrasonically assisted nitration reactions (USNR). Cavitation induces very high local temperatures and the liquid and enhanced mass transfer. **Figure 2** represents the development and collapse of bubbles due to cavitation.

### 3.3. Recyclisation of the Catalyst

For large scale application of the reaction, it would be desirable to recycle the catalyst. The reaction with 0.25 equivalents or 0.5 equivalents of ammonium molybdate is very slow. Therefore the reaction is carried out at 1:1:1 molar ratio. Furthermore, the catalyst is readily recycled by a simple filtration after ensuring that it is free from nitrate ions by washing with water. The recovered catalyst was used as such and found to be active for further nitrations

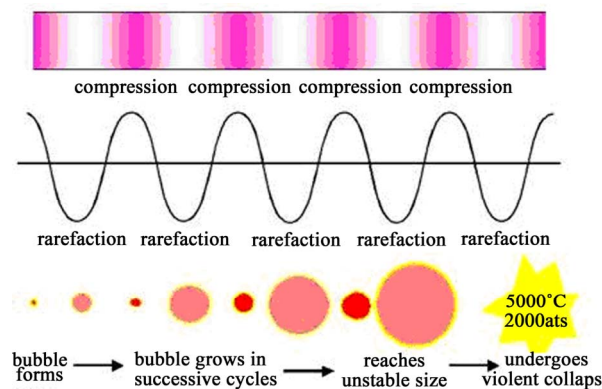


Figure 2. The process of acoustic cavitation.

with no reduction in yield and with no change in selectivity. Benzaldehyde, benzamide and benzoic acid were selected as the substrates to examine the catalyst recycling. To ascertain how long the efficacy of the catalyst would last, the reaction was repeated minimum three to four times using the same batch of catalyst. The results were reproducible with good accuracy.

#### 4. Conclusion

In conclusion we have demonstrated that ultrasonically assisted nitration reactions (USANR) underwent smoothly with good yields and high regioselectivity is observed for anilides, moderately and non-activated aromatic compounds. Longer reaction times (6 - 8 hrs) of metal catalyzed reactions reduced to (1 - 2 hrs) under sonication. When ortho position is blocked para derivatives are obtained while ortho nitro products are obtained when para position is blocked. In case of USANR of aromatic carbonyl and related compounds the effect of sonication is highly significant; this could be seen from the observed reaction times ranging in few minutes.

#### 5. Acknowledgements

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## Supplementary Data

P1112-1008 N-0041

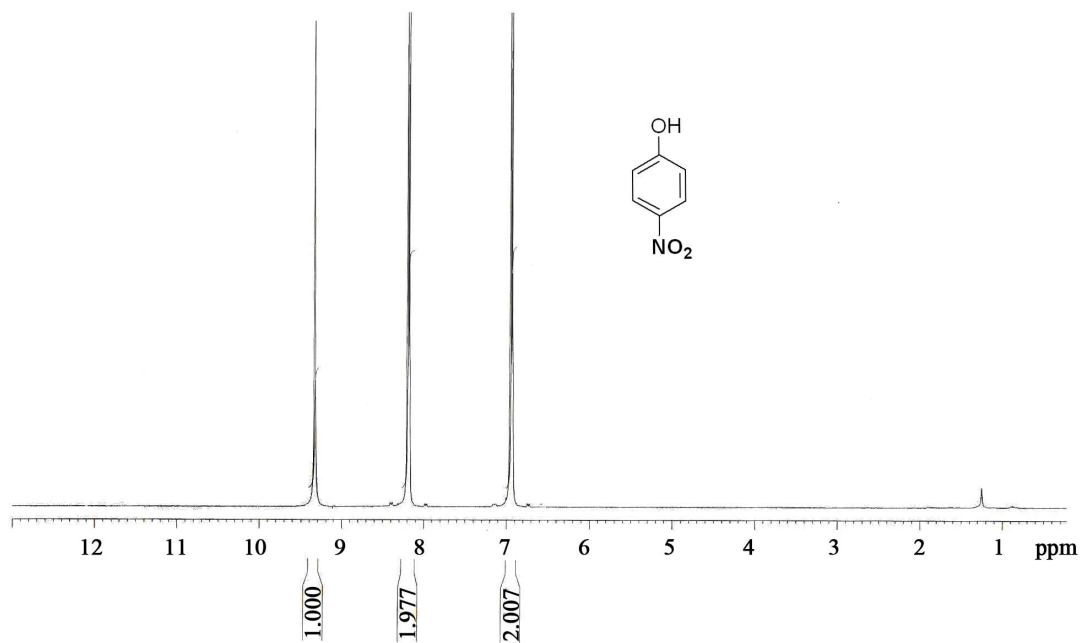


Figure S1. HNMR spectrum of 4-nitro phenol.

P1112-1008 N-008

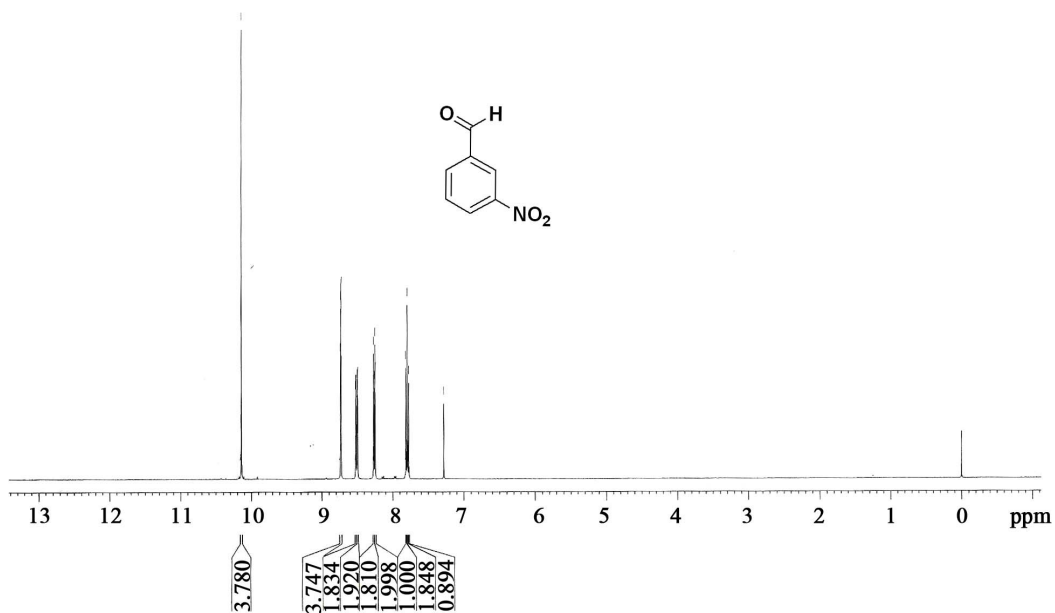


Figure S2. HNMR spectrum of 3-nitro benzaldehyde.

P1112-1008 N-009

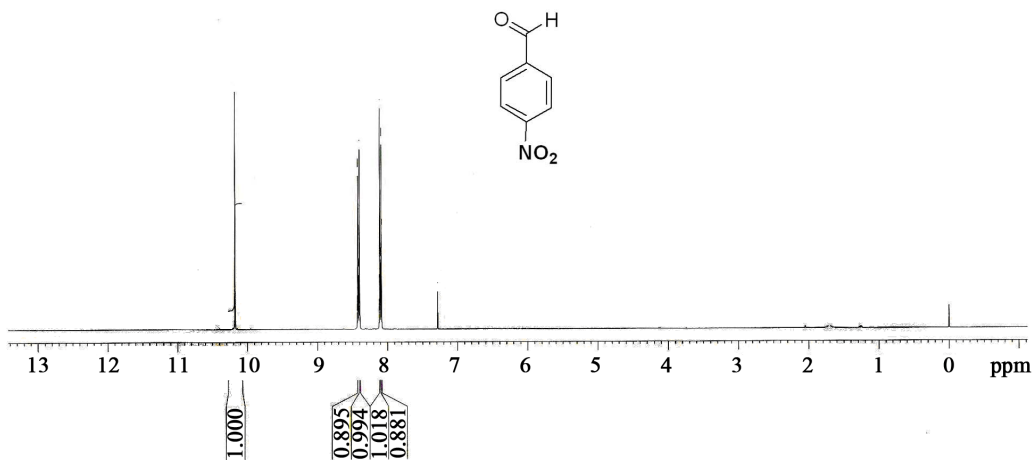


Figure S3. HNMR spectrum of 4-nitro benzaldehyde.

P1112-1008 N-020

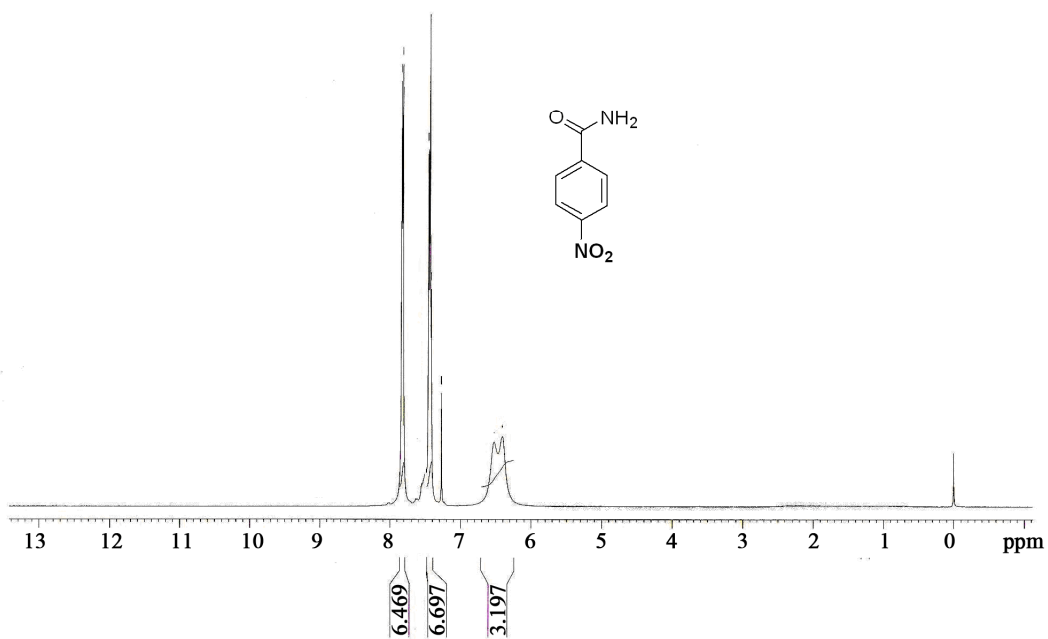


Figure S4. HNMR spectrum of 4-nitro benzamide.

P1112-1008 N-023

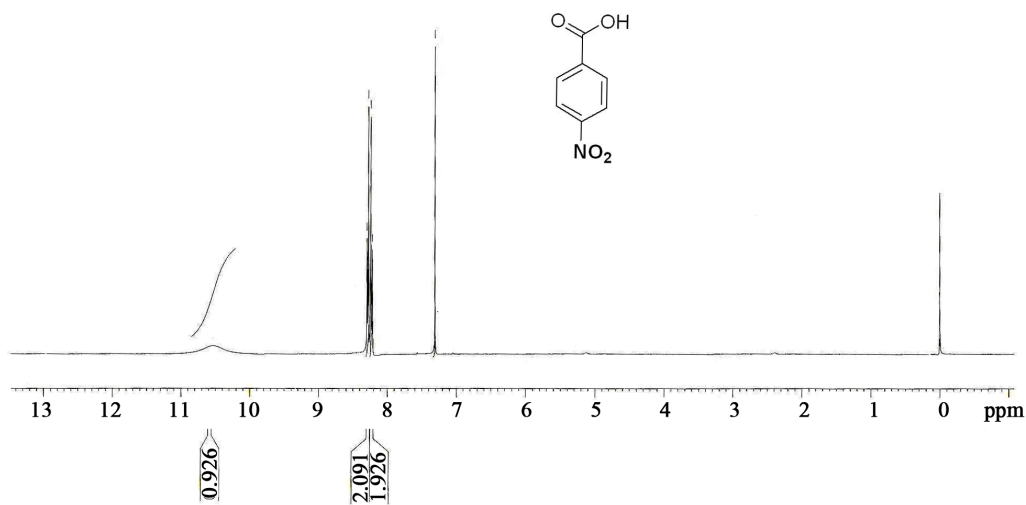


Figure S5. HNMR spectrum of 4-nitro benzoic acid.

P1112-1008 N-001

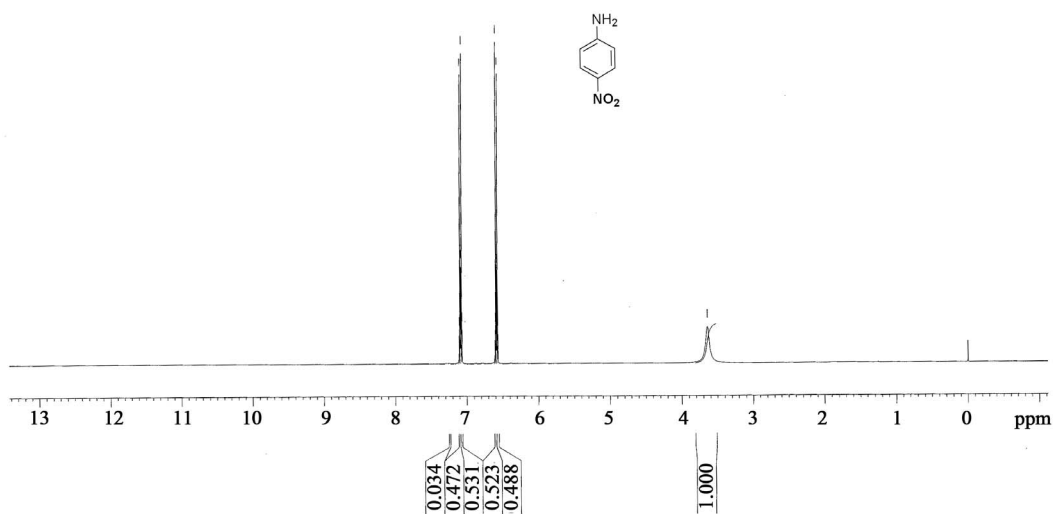


Figure S6. HNMR spectrum of 4-nitro aniline.

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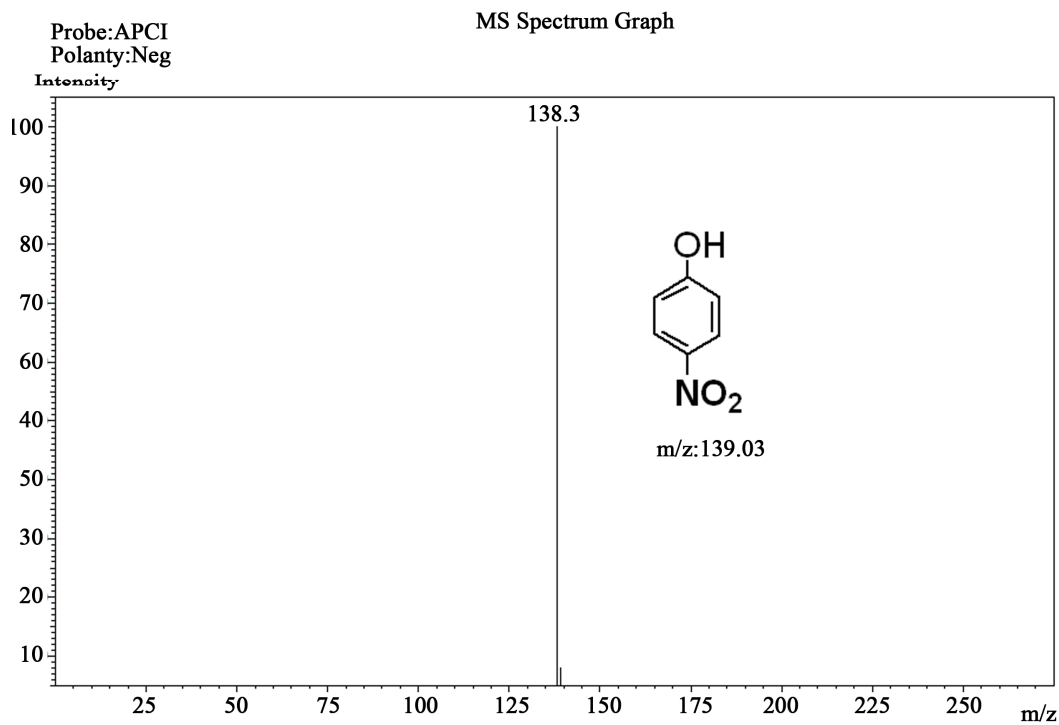


Figure S7. Mass spectrum of 4-nitro phenol.

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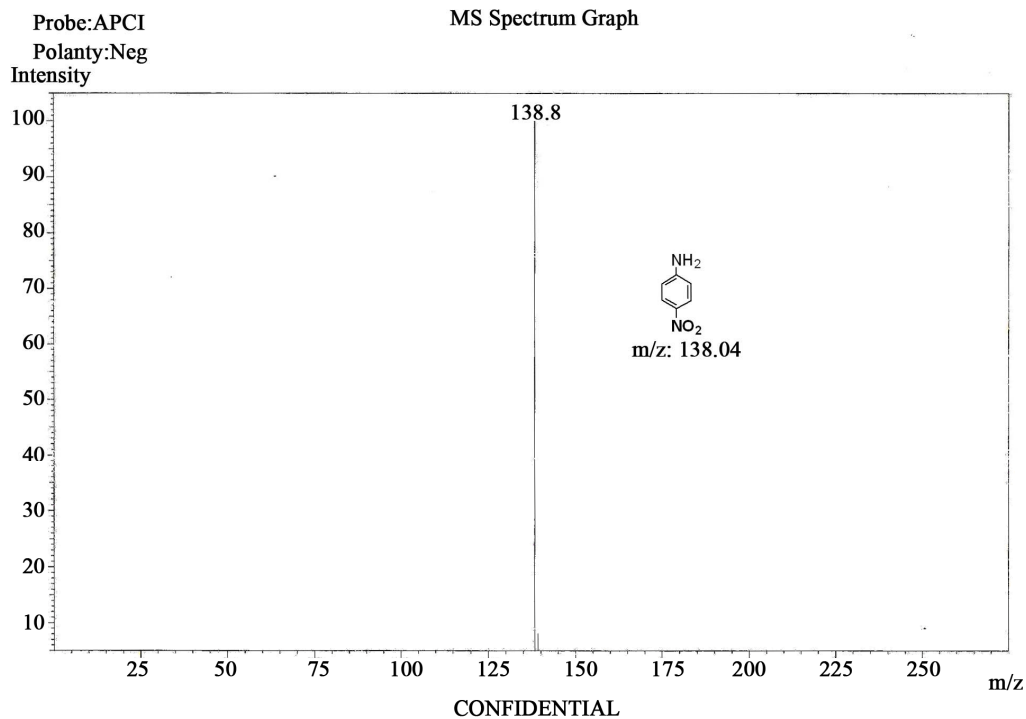


Figure S8. Mass spectrum of 4-nitro aniline.

## SAPALA ORGANICS PVT LTD

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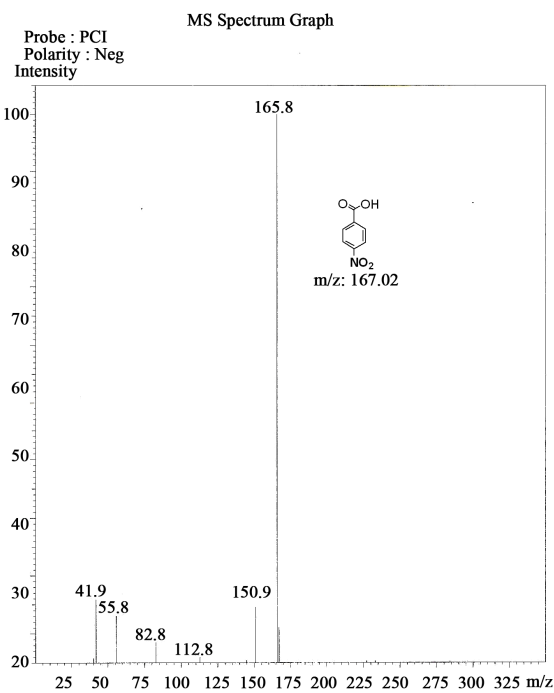


Figure S9. Mass spectrum of 4-nitro benzoic acid.