PULSE RADIOLYSIS STUDIES OF AZA ANALOGUES OF NUCLEIC ACID COMFONENTS M. Faraggi and I. Rosenthal

The reaction of hydrated electron and OH radical with the anticancer drug  $\sigma$ -azacytidine and other aza-nucleic acid analogues was measured. It was found that the reaction of both radicals is fast, approaching the diffusion controlled process. The transient spectra of the radicals resulting from the reactions of OH radicals and  $e_{aq}$  were determined. Normally these spectra exhibited maxima at  $\lambda \sim 285$  nm and a shoulder at  $\sim 350$  nm.

PROPERTIES OF CUPRIC-HYDRIDE FORMED IN THE REACTION OF AQUA CUPROUS IONS WITH HYDROGEN ATOMS. A PULSE-RADIOLYTIC STUDY<sup>(1)</sup> W. A. Mulac<sup>\*</sup> and D. Meyerstein

The spectrum of  $\operatorname{Cu-H}_{aq}^+$  formed via  $\operatorname{Cu}^+ + \operatorname{H} \rightarrow \operatorname{Cu-H}_{aq}^+$  in aqueous solutions was determined. In neutral solutions it decomposes via  $\operatorname{Cu} \operatorname{H}_{aq}^+ + \operatorname{H}_2 0 \rightarrow \operatorname{Cu}_{aq}^{2+} + \operatorname{H}_2 + \operatorname{OH}_{aq}^-$  with a rate of  $4 \times 10^3 \, \mathrm{s}^{-1}$ . The properties of  $\operatorname{Cu-H}_{aq}^+$  were considered and compared with those of similar  $\operatorname{Cu}^{II} - \operatorname{R}$  compounds.

## **REFERENCE:**

1. Mulac, W. A. and Meyerstein, D., Inorg. Chem., in press.

REACTIONS OF  $B_{12r}$  WITH ALIPHATIC FREE RADICALS. A PULSE RADIOLYSIS STUDY<sup>(1)</sup> W. A. Mulac<sup>\*</sup> and D. Meyerstein

The spectra of the intermediates formed in the reactions of  $B_{12r}$ with the free radicals,  $Br_2$ ,  $CO_2$ ,  $CH_2C(CH_3)_2OH$ ,  $C(CH_3)_2OH$ ,  $CH_2CHO$ and  $CH(OH)CH_2OH$  were determined. The results indicate that  $Br_2$ oxidizes  $B_{12r}$  to  $B_{12a}$  via an inner sphere mechanism, and  $CO_2$  reduces  $B_{12r}$  to  $B_{12s}$ . All the aliphatic free radicals studied, R, react with  $B_{12r}$  yielding as the first product a pseudo coenzyme denoted  $Co^{III}-R$ .  $Co^{III}-CH_2C(CH_3)_2OH$  is stable for over a second in the pH range 3-10 as is  $Co^{III}-CH_2CHO$ . The latter compound hydrolyzes in acid solutions to yield  $B_{12a}$  and  $CH_3CHO$ .  $Co^{III}-C(CH_3)_2OH$  and  $Co^{III}-CH(OH)CH_2OH$  decompose

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