

Synthesis and Comparative Antibacterial Activity of Fatty Acid Capped Silver Nanoparticles

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Abstract

In the present work, the chemical reduction method has been used to synthesize silver nanoparticles using stearic acid capping agents in different concentrations. These nanoparticles are characterized by ultraviolet-visible spectroscopy and dynamic light scattering (DLS). The size of the synthesized silver nanoparticles found between 80-100 nm and stable up to 5 months. These nanoparticles show a very good bactericidal influence on *E. coli* and *S. aureus*. In this method silver nanoparticles have synthesized at room temperature without using any inert atmosphere, this is the advantage of this method over others.

Keywords: Silver nanoparticles, Stearic acid, UV-Vis. spectroscopy, Antibacterial

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(Received: April 12, 2020; accepted: August 26, 2020)

Citation: Sharma R. Synthesis and Comparative Antibacterial Activity of Fatty Acid Capped Silver Nanoparticles. *J Pure Appl Microbiol.* 2020;14(3):1941-1947. doi: 10.22207/JPAM.14.3.33

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INTRODUCTION

Nanotechnology, at this time, is a prominent field which is used for the production of nanomaterials and their utilisation¹Persimmon, Ginkgo, Magnolia and Platanus. Chemical reduction method to synthesize silver nanoparticles reduces atomic silver (Ag^+) of silver nitrate to zero-valent silver (Ag^0) i.e. nanosilver. Silver nanoparticles having 1 nm and 100 nm size have good medicinal properties². Silver compounds are very well known antimicrobial agents from ancient times. These are toxic also for several bacteria^{3,4}. Because of these applications, silver is frequently used and has an important role in the medical field.

Silver nanoparticles are more and more used in different areas, together with health care, food, and manufacturing purposes, because of their exclusive properties. These contain visual, thermal, high electrical conductivity, and biological properties. These properties of silver nanoparticles make them useful as biological sensors. Silver nanoparticles are used as an antibacterial, antiviral, antifungal coating⁵. Nowadays biomedical devices containing silver nanoparticles, liberate low levels of silver ions continuously to give protection from bacteria.

Silver nanoparticles can be reduced by reducing agents like sodium borohydride in the presence of fatty acids like myristic acid at room temperature in a medium containing methanol-water. Fatty acids are long-chain lipids with hydrocarbon compounds and carboxylic groups on the terminal. It has been investigated that pH increases the visual properties of silver nanoparticles. The silver nanoparticle that is capped by fatty acid namely myristic acid are 3 to 7 nanometer in size and has stability for several months. This way of synthesis of myristic acid could be beneficial more than the last information of carboxylate stabilize silver nanoparticles. The benefits of this method include the production of silver nanoparticles at room temperature in aqueous medium⁶. Das et al.⁷ prepared the silver nanoparticle by using ethanol in the presence of sodium linoleate through the reduction method. They reported that nanoparticle stops the growth of microorganisms like bacteria, fungi, parasites, etc.

Rao and Trivedi prepared some fatty acid stabilized silver nanoparticles by biphasic synthesis

using three reducing agents like triethanolamine, formaldehyde, and dimethyl amide⁸. They found the role of experimental condition on the size of particles. It has been found that as the fatty acid chain length is increased, the stability of silver nanoparticles is also increased. The several experimental conditions like concentration and temperature inclined the dimension of silver nanoparticles frequently.

Cinar et al. prepared silver nanoparticles by oleylamine and reduction of oleic acid. They reported the use of these nanoparticles for the manufacturing of nano cable through coaxial electrospinning⁹. Different fatty acid capped silver nanoparticles were prepared by different methods¹⁰⁻¹⁶. Silver nanoparticles have good applications in different fields¹⁷⁻²².

Industrial applications of silver nanoparticles are very limited because they get easily oxidized during formation. To overcome such type of problems, fatty acids and their compounds have been used as stabilizers to synthesize metal nanoparticles²³. Generally, fatty acids are well known to coordinate with metal nanoparticles^{24,25}.

EXPERIMENTAL

MATERIALS

All the chemicals used in the experiment were from Central Drug House, India. Triple distilled water is used for making all the solutions. Analytical reagent (AR) grade reagents and solvents are used to perform this work.

Synthesis of silver nanoparticles (AgNPs)

In this work silver nitrate (AgNO_3) was reduced to get AgNPs using hydrazine hydrate and stearic acid aqueous solution¹². A solution of stearic acid was prepared by different concentrations of stearic acid (1 g, 0.6 g, 0.3 g). Each of these concentrations of stearic acid was dissolved in 50 ml triple distilled water separately and stirred for 10 minutes at 80°C to obtain a clear solution. Then add 1.2 ml of liquor NH_3 in this solution and stir this solution by using a magnetic stirrer for about 10 minutes. After stirring add 0.5 grams of silver nitrate, stir for about 5 minutes, and add 4 ml hydrazine hydrate dropwise. Stirring was continued for 1 hour to get a transparent solution. Then centrifugation was done so that supernatant and pallet of silver nanoparticles separated.

Physicochemical characterization

Initially, the synthesis of AgNPs (diluted 10 times with triple distilled water) was established by absorption spectra at the wavelength of 200–600 nm. The automatic mode was followed to determine the size of AgNPs with the technique Dynamic light scattering using instrument Zetasizer Nano ZS (Malvern Instruments, UK).

Antibacterial Activity

Preparation of Media

The standard formulation given in Bacteriology Manual²⁶ was followed for the preparation of all media.

Composition

- peptone - 0.5%
- NaCl - 0.5%
- agar -1.5%
- beef extract - 0.3%
- distilled water
- pH - (6.8) at room temperature.

All the above ingredients were completely dissolved in triple distilled water. This medium was taken into a flask and this flask was autoclaved for 30 minutes at 121°C for sterilization. 15 ml of this sterilized medium was poured in sterilized Petri plates aseptically and incubated for about 24 hours at 37°C.

Antibacterial activities of AgNPs were performed by the disc diffusion technique. The prepared AgNPs were tested against the antibacterial resistance *E. coli* (gram –ve) and *S. aureus* (gram +ve). 10 µl of *E. coli* and 10 µl of *S. aureus* were spread into the nutrient agar medium plates. The swab was used to spread the smear evenly.

Sterile discs were prepared by dipping in 1 mM AgNO₃ solution, nano colloidal solution synthesized by 0.3 g, 0.6 g, and 1 g stearic acid. The so prepared 4 sterile discs were placed in the *E. coli* culture plate and 4 were placed in the *S. aureus* culture plate to observe the zone of inhibition. All culture plates were gestated for about 24 hours at 37° C and the region of inhibition was measured in mm.

RESULTS AND DISCUSSION

Synthesis of silver nanoparticles

The color change of the solution from light brown to dark brown (Fig. 1) approves the synthesis of AgNPs and the absorption band at

around 420 nm also confirms the synthesis of AgNPs in the suspension.

Physicochemical characterization

Some factors like concentration affect the synthesis of small-sized monodispersed nanoparticles. So it is necessary to optimize different concentrations of stearic acid solutions for the synthesis of AgNPs. The synthesis of AgNPs



Fig. 1. AgNPs

at a concentration of 0.3g/50ml showed maximum absorbance. The reduction was continued for 60 minutes and a characteristic surface plasma resonance band was noticed around 420 nm, which confirms the formation of AgNPs (Fig. 2).

Metal nanoparticles having surface plasma resonance (SPR) can easily be characterized by absorption spectra. The SPR takes place

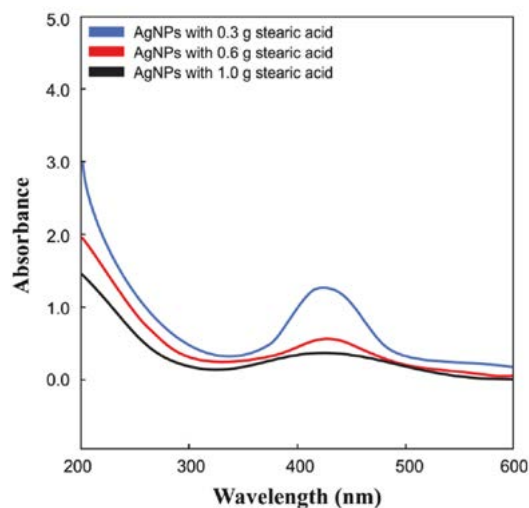


Fig. 2. UV-Vis spectra

because of the fluctuation of the electrons present on the surface of metal nanoparticles due to the external energy sources. By this specific SPR peak, the physical nature of nanoparticles can be studied

very easily. SPR mainly depends upon the size of nanoparticles and the properties of dielectric medium²⁷⁻²⁹. The broadband shows broad size and aggregation of nanoparticles while a narrow band

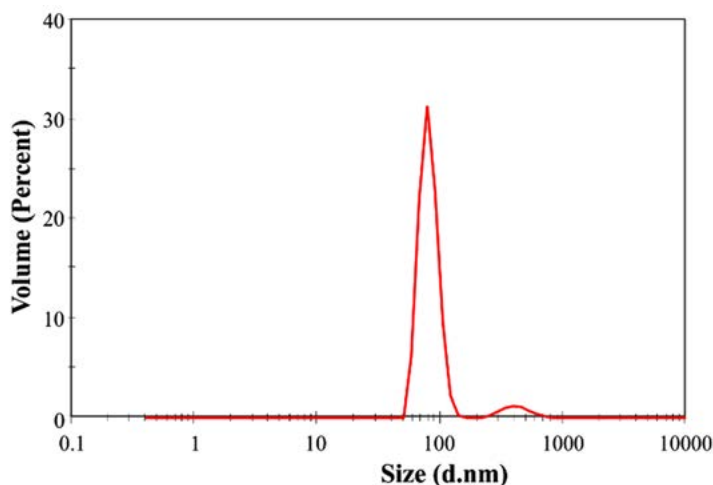


Fig. 3. Particle size analysis

denotes monodispersed particles³⁰. A sharp band represents the narrow distribution of the AgNPs³¹. The position and number of peaks also change with the shape and size of nanoparticles.

Spherical nanoparticles show a particular SPR band and anisotropic particles show more than two SPR bands³². The size of the synthesized

nanoparticles was analyzed by dynamic light scattering. In the DLS technique, the size is measured by illuminating the nanoparticles in Brownian motion by a laser beam. The light scattered by the nanoparticles is then analyzed by the auto-correlator. The synthesized AgNPs were 80-100 nm in size (Fig. 3).

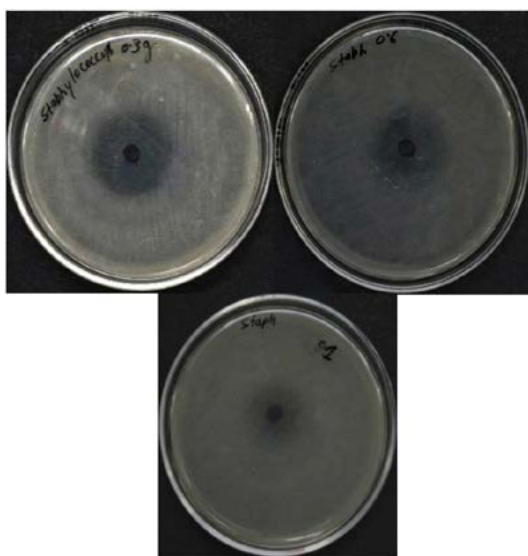


Fig. 4(a). Zone of inhibition against *S. aureus*

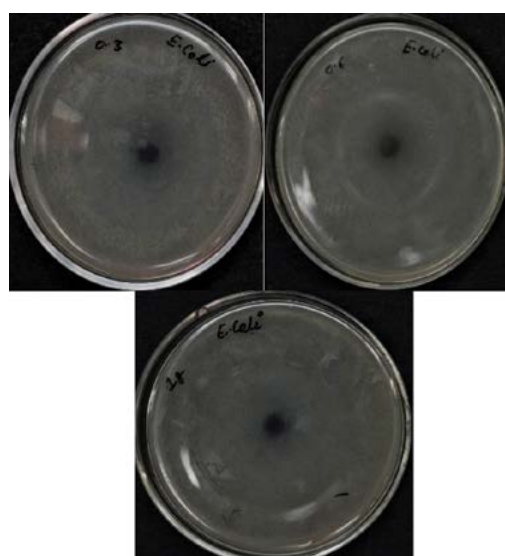


Fig. 4(b). The zone of inhibition against *E. coli*

Table 1. Zone of inhibition

AgNPs with different concen. of stearic acid (in gram)	<i>E. coli</i> (in mm)	<i>S. aureus</i> (in mm)
0.3	20	6
0.6	12	4
1	6	2
1 mM AgNO ₃	1	1

Antibacterial activity

Larger surface area is favorable for bacterial surface interactions that are the reason silver nanoparticles show better antibacterial activities than silver ions. As the particle size decreases (nanoscale), surface area increases which facilitate bacterial action³³. The zone of inhibition obtained against *E. coli* and *S. aureus* clearly mentions the influential nature of synthesized AgNPs (Table 1). A noticeable difference in the zone of inhibition was observed for each solution having a different amount of stearic acid as shown in Figs 4a and 4b. Different researchers reported a number of mechanisms for the antibacterial action of AgNPs³⁴⁻³⁶.

The stability of silver nanoparticles

Silver nanoparticles have many applications because of their stability. If silver nanoparticles so obtained show precipitation after storage, the number of applications will be reduced. The silver nanoparticles colloid fresh and after storage of 5 months are shown in Fig. 5. No precipitation was noticed in the colloidal solution of so obtained stearic acid stabilized silver nanoparticles even after storage for a long time, so it becomes clear that stearic acid is a good stabilizing agent to impart stability to silver nanoparticles. Precipitation and oxidation of silver nanoparticles are prevented by the high surface activity, involving organic functional groups of stearic acid.

CONCLUSION

Hydrazine hydrate and stearic acid are used as a reducing agent to reduce silver nitrate to silver nanoparticles. It is successfully established that stearic acid is a good coating agent to synthesize AgNPs in a suspension. The so prepared

**Fig. 5.** AgNPs (a) fresh (b) after 5 months

silver nanoparticles showed good stability and no precipitation was observed after storage of 5 months and longer. It is clear from the zone of inhibition that the concentration affects the zone of inhibition as the concentration of capping agent i.e. stearic acid increased, the zone of inhibition decreased.

ACKNOWLEDGMENTS

Author is thankful to the Director, IAH, G.L.A. University, Mathura for financial support of this research.

FUNDING

None.

ETHICS STATEMENT

Not applicable.

DATA AVAILABILITY

All the analysis related to this work are incorporated in the manuscript.

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